

Roll No. of Candidate : _____

CHEMISTRY

Intermediate Part-I, Class 11th (1stA 323- II)

Paper : I

Group – II

Time: 20 Minutes

OBJECTIVE Code : 6484

Marks: 17

Note: You have four choices for each objective type question as A, B, C and D. The choice which you think is correct, fill that circle in front of that question number. Use marker or pen to fill the circles. Cutting or filling two or more circles will result in zero mark in that question.

1. 1 - Pressure remaining constant, at which temperature the volume of a gas will become twice of what it is at 0° C
(A) 546° C (B) 200° C (C) 546 K (D) 273 K
- 2 - An aqueous solution of ethanol in water may have vapour pressure
(A) equal to that of water (B) equal to that of ethanol
(C) more than that of water (D) less than that of water
- 3 - In zero order reaction, the rate is independent of
(A) temperature of reaction (B) concentration of reactants
(C) concentration of products (D) none of these
- 4 - Liquid Hydrocarbon is
(A) Methane (B) Pentane (C) Hexane (D) Propane
- 5 - If a strip of Cu metal is placed in a solution of FeSO₄
(A) Cu will be deposited (B) Fe is precipitated out
(C) Cu and Fe both dissolve (D) no reaction takes place
- 6 - The value of pH of pure water at 25° C is
(A) 14 (B) 7 (C) 1×10^{-14} (D) 1×10^{14}
- 7 - When 6 d orbital is complete, the entering electron goes in to
(A) 7 f (B) 7 p (C) 7 s (D) 7 d
- 8 - Calori is equal to
(A) 0.4184 J (B) 41.84 J (C) 4.184 J (D) 418.4 J
- 9 - The oxidation No. of Nitrogen in HNO₃ is
(A) +3 (B) -3 (C) -5 (D) +5
- 10 - During the process of crystallization the hot saturated solution
(A) is cooled very slowly to get large size crystals
(B) is cooled at a moderate rate to get medium size crystals
(C) is evaporated to get the crystals of the product
(D) is mixed with immisible to get the pure crystals of the product
- 11 - The mass of one mole of electron is
(A) 1.008 mg (B) 0.55 mg (C) 0.184 mg (D) 1.673 mg
- 12 - Feeling uncomfortable breathing in unpressurized cabin is due to
(A) high pressure of CO₂ (B) low pressure of CO₂
(C) low pressure of O₂ (D) high pressure of O₂
- 13 - The change in heat energy of a chemical reaction at constant temperature and pressure is called
(A) enthalpy change (B) heat of sublimation
(C) bond energy (D) internal energy change
- 14 - The pH of 10^{-3} mol dm⁻³ of an aqueous solution of H₂SO₄ is
(A) 3.0 (B) 2.7 (C) 2.0 (D) 1.5
- 15 - Which of the following species has unpaired electrons in the antibonding molecular orbitals ?
(A) O₂²⁺ (B) N₂²⁻ (C) B (D) F₂
- 16 - Diamond is a bad conductor because
(A) it has a tight structure
(B) there are no free electrons present in the crystal of diamond to conduct electricity
(C) it has a heigh density
(D) is transparent to light
- 17 - 27 g of Al will react how much mass of O₂ to produce Al₂O₃
(A) 8 g of Oxygen (B) 16 g of Oxygen (C) 32 g of Oxygen (D) 24 g of Oxygen